GSC International
Phone: 417.374.7431
Fax: 417.374.7442
Toll Free: 888.756.4592
service@gosciencecrazy.com
2076 James River Court
Nixa, Missouri 65714
GO Science Crazy

## Deluxe Boyle's Law Apparatus \#BOYLAW-DLX



## Introduction

This apparatus is useful to study the basics of Boyle's Law and to estimate the atmospheric pressure. It becomes a simple way to approach the experimental side of physics and to introduce at sight the concepts of vacuum, pressure, density, and more. With the tube filled with a heavy oil (such as motor oil) and connected to a closed reservoir, this apparatus allows you to study the aeriform substance in the expansion container. By raising or lowering the other tube end, compression or rarefaction of the gas is obtained.

## $\mathbf{P V}=k$

Boyle's law states that the volume $\mathbf{V}$ of gas molecules is inversely proportional to its pressure $\mathbf{P}$. Therefore:


The variation of the height of the oil column implies a simultaneous change of the volume occupied by the substance.

## How to Use

1. Place the apparatus vertically on a balanced and smooth surface.
2. Place the open-ended glass tube on the upper side and the closed-ended tube on the lower side.
3. Now place the funnel on the open-ended glass tube and fill the glass tube with motor oil (or some other heavy oil) until the oil level reaches the halfway point of the closed-ended glass tube.
4. Fix the slider holding the closed-ended glass tube at the lowest point of the scale and note down the coinciding point of the top of the closed-ended glass tube.
5. Now slide the open-ended glass tube by just freeing the knob provided with the slider and moving it downward.
6. Note down the height of empty space inside the open-ended glass tube and the oil level readings of the closedend glass tube at different heights.
7. Take at least 8 readings at different points.
8. Draw a between pressure P taking it along X -axis and volume V taking it along Y -axis.

## Observation Table

| Sr. No. | Height of the <br> empty space in <br> open tube <br> pressure (P) | Height of the <br> mercury level in <br> the closed tube <br> (h1) | Top point reading <br> of the closed tube <br> (h2) | Height of the <br> empty space in <br> the closed tube <br> (volume V = h2 - <br> h1) |
| :---: | :---: | :---: | :---: | :---: |
| 1 |  |  |  |  |
| 2 |  |  |  |  |
| 3 |  |  |  |  |
| 4 |  |  |  |  |
| 5 |  |  |  |  |
| 6 |  |  |  |  |
| 7 |  |  |  |  |
| 8 |  |  |  |  |

