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Antioxidant effects of reduced water produced by electrolysis of sodium chloride solutions

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Abstract

Antioxidant vitamins and enzymes such as superoxide dismutase, catalase and glutathione peroxidase are considered to function as scavengers against reactive oxygen species and to provide protection against reactive oxygen species, including free radicals. Although antioxidants such as L-ascorbic acid, *d*-catechin and quercetin dehydrate show superoxide dismutation activity, using reduced water produced in the cathode side by electrolysis as a solvent instead of 2 mM NaC1 solution of the same pH level as the reduced water increased the superoxide dismutation activity of these antioxidants. Moreover, neither the reduced water nor its electrolyte solution showed any superoxide dismutation activity by itself. On the other hand, the reduced water was able to decrease hydrogen peroxide levels. It has been found that the behaviour of H_2 in reduced water, which was activated by a platinum electrode, differed from that of H_2 introduced by bubbling of hydrogen gas. The former decreased H_2O_2 , whereas the latter did not. These results suggest strongly that the increase in superoxide dismutation activity, with a proton donor such as L-ascorbic acid, is due to an increase in the dissociation activity of water while the scavenging activity for H_2O_2 is due to activated dissolved H_2 in the reduced water.

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antioxidant dismutation activity

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reduced water superoxide radicals