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$$b = \frac{Z e^2 \cot \theta / 2}{4\pi \epsilon_0 \left(\frac{1}{2} m v^2 \right)}$$

Trace the Mind Map
► First Level ► Second Level ► Third Level

- In 1898, J.J. Thomson proposed the first model of atom known as plum-pudding model.
- In 1911, Rutherford prepared planetary model of atom.
- In 1913, Niels Bohr prepared a model of Hydrogen atom based on quantum theory.

- Hydrogen gas heated in a sealed tube emits radiation when passed through prism components of different wavelength appear.
- Wavelength in each series given by

$$\nu = \frac{1}{\lambda} = R \left[\frac{1}{n_i^2} - \frac{1}{n_f^2} \right]$$

$n_f > n_i$
Lyman series [U.V. region]

$n_i = 1, n_f = 2, 3, 4, \dots$
 $\lambda_{min} = 912 \text{ \AA}$, $\lambda_{max} = 1216 \text{ \AA}$

Balmer series [Visible region]
 $n_i = 2$, $n_f = 3, 4, 5, \dots$
 $\lambda_{min} = 3648 \text{ \AA}$, $\lambda_{max} = 6563 \text{ \AA}$

Paschen series [IR region]
 $n_i = 3$, $n_f = 4, 5, 6, \dots$
 $\lambda_{min} = 8208 \text{ \AA}$, $\lambda_{max} = 18761 \text{ \AA}$

Brackett series [IR region]
 $n_i = 4$, $n_f = 5, 6, 7, \dots$
 $\lambda_{min} = 14592 \text{ \AA}$, $\lambda_{max} = 40533 \text{ \AA}$

Pfund [IR region]
 $n_i = 5$, $n_f = 6, 7, 8, \dots$
 $\lambda_{min} = 23850 \text{ \AA}$, $\lambda_{max} = 74618 \text{ \AA}$

- Energy of electron in each stationary orbits is

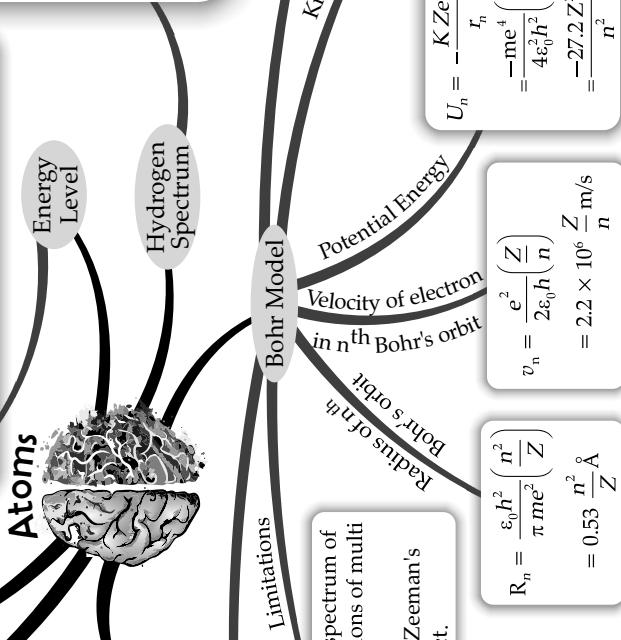
$$\text{Given by } E_n = -13.6 \frac{Z^2}{n^2} \text{ eV}$$

Where $n = 1, 2, 3, \dots$

Z = atomic number of atom

- These stationary energy orbits are also called energy levels.
- When electron jumps from higher energy level to lower energy level, it releases energy.

$$\Delta E = E_f - E_i = 13.6 \left[\frac{1}{n_i^2} - \frac{1}{n_f^2} \right] Z^2 \text{ eV}$$



- Atoms have a central, massive, positively charged core called nucleus around which electrons revolve.
- Size of nucleus ≈ 1 fermi $= 10^{-15} \text{ m}$

- Electron revolves around the nucleus in stationary orbits.
 - Angular momentum of electron $mvT_n = n \times \frac{h}{2\pi}$ where, n is an integer
- It is also known as principal Quantum Number.
- It explains spectra of hydrogen or hydrogen like $[\text{He}^+, \text{Li}^{++}]$ atoms.

