



Acids & Bases

Part 1

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Summary of All Formulas

[H₃O⁺] to pH

$$\text{pH} = -\log [\text{H}_3\text{O}^+]$$

pH to [H₃O⁺]

$$[\text{H}_3\text{O}^+] = 10^{-\text{pH}}$$

[OH⁻] to pOH

$$\text{pOH} = -\log [\text{OH}^-]$$

pOH to [OH⁻]

$$[\text{OH}^-] = 10^{-\text{pOH}}$$

pH to pOH
or
pOH to pH

$$\text{pH} + \text{pOH} = 14$$

[H₃O⁺] to [OH⁻]
or
[OH⁻] to [H₃O⁺]

$$[\text{H}_3\text{O}^+] \times [\text{OH}^-] = 1.0 \times 10^{-14}$$

[OH⁻] to pH

- $\text{pOH} = -\log [\text{OH}^-]$

- $\text{pH} + \text{pOH} = 14$

pH to [OH⁻]

- $\text{pH} + \text{pOH} = 14$

- $[\text{OH}^-] = 10^{-\text{pOH}}$